MATERIAL SAFETY DATA SHEET

SRM Supplier: National Institute of Standards and Technology
Standard Reference Materials Program
Bldg. 202 Rm. 211
Gaithersburg, Maryland 20899

SRM Number: 3158
MSDS Number: 3158
SRM Name: Thallium Standard Solution
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MSDS Coordinator: Joylene W.L. Thomas
Phone: (301) 975-6776
FAX: (301) 926-4751
e-mail: SRMMSDS@nist.gov

SECTION I. MATERIAL IDENTIFICATION

Material Name:: Thallium Standard Solution

Description: SRM 3158 is a single element solution prepared gravimetrically to contain a nominal 10 mg/g of thallium with a nitric acid volume fraction of 10 %.

Other Designations: **Thallium** in **Nitric Acid** (aqua fortis; hydrogen nitrate; azotic acid; engravers acid); ***Thallium Nitrate** (thallium (I) nitrate; thallous nitrate; thallium mononitrate) in **Spectrometric Standard Solution**

Name	Chemical Formula	CAS Registry Number
Nitric Acid	HNO_3	7697-37-2
Thallium Nitrate	$TlNO_3$	10102-45-1
Thallium	Tl	7440-28-0

DOT Classification: Nitric Acid, Solution, UN2031

Manufacturer/Supplier: It is available from a number of suppliers.

*The addition of thallium to nitric acid, along with other intermediate chemical reactions, forms thallium nitrate which will precipitate upon evaporation or drying of the solution.

SECTION II. HAZARDOUS INGREDIENTS

Hazardous Components	Nominal Concentration (%)	Exposure Limits and Toxicity Data
Nitric Acid	10	ACGIH TLV-TWA: 2 mg/kg or 5 mg/m^3
		OSHA TLV-TWA: 2 mg/kg or 5 mg/m^3
		Human, Oral: LD_{LO}: 430 mg/kg
Thallium Nitrate	1.91	ACGIH TLV-TWA: 0.1 mg/m^3
		OSHA TLV-TWA: 0.1 mg/m^3
		Rat, Intraperitoneal: LD_{50}: 21 µg/kg
		Man, Oral: TD_{LO}: 73 mg/kg
Thallium	1	ACGIH TLV-TWA: 0.1 mg/m^3
		OSHA TLV-TWA: 0.1 mg/m^3
		Man, Oral: TD_{LO}: 5714 µg/kg

SECTION III. PHYSICAL/CHEMICAL CHARACTERISTICS

Nitric Acid	Thallium Nitrate	Thallium
Appearance and Odor: a white to slightly yellow liquid that darkens to a brownish color upon aging and exposure to light; pungent odor	**Appearance and Odor:** colorless to white crystals; odorless	**Appearance and Odor:** white, odorless crystals
Relative Molecular Mass: 63.02	**Relative Molecular Mass:** 266.39	**Relative Atomic Mass:** 204.383
Density: 1.0543 (10% nitric acid)	**Density:** 5.55	**Density:** 11.85
Solubility in Water: soluble	**Solubility in Water:** slightly soluble	**Solubility in Water:** insoluble
Solvent Solubility: decomposes in alcohol	**Solvent Solubility:** soluble in acetone	**Solvent Solubility:** soluble in nitric acid and sulfuric acids

NOTE: The physical and chemical data provided are for the pure components. Physical and chemical data for this thallium/nitric acid solution do not exist. The actual behavior of the solution may differ from the individual components.

SECTION IV. FIRE AND EXPLOSION HAZARD DATA

Flash Point: N/A **Method Used:** N/A **Autoignition Temperature:** N/A

Flammability Limits in Air (Volume %): **UPPER:** N/A
LOWER: N/A

Unusual Fire and Explosion Hazards: Although nitric acid does not burn, it is a powerful oxidizing agent that can react with combustible materials to cause fires. Thallium and thallium nitrate are negligible fire hazards when exposed to heat or flames.

Extinguishing Media: Use extinguishing media that is appropriate to the surrounding fire. Use a water spray to dilute nitric acid and to absorb liberated oxides of nitrogen.

Special Fire Procedures: Fire fighters should wear a self-contained breathing apparatus (SCBA) with a full face piece in the pressure demand or positive mode and other protective clothing.

SECTION V. REACTIVITY DATA

Stability: **__X__ Stable** **_____ Unstable**

Conditions to Avoid: Avoid contact with combustible and other incompatible materials.

Incompatibility (Materials to Avoid): Keep nitric acid away from organic materials, plastics, rubber, and some forms of coatings. Nitric acid is incompatible with chlorine and metal ferrocyanide. Thallium is incompatible with halogens. Thallium nitrate should be kept from combustible materials and reducing agents.

See Section IV: *Unusual Fire and Explosion Hazards*

Hazardous Decomposition or Byproducts: Hazardous decomposition of nitric acid and/or thallium nitrate can produce various nitrogen oxides, including nitric oxide (NO), nitrogen dioxide (NO_2), nitrous oxide (N_2O), as well as nitric acid mist or vapor. Thermal decomposition of thallium may release toxic and/or hazardous gases.

Hazardous Polymerization: **_____ Will Occur** **__X__ Will Not Occur**

SECTION VI. HEALTH HAZARD DATA

Route of Entry: __X__ **Inhalation** __X__ **Skin** __X__ **Ingestion**

Health Hazards (Acute and Chronic): Nitric Acid: Nitric acid may be fatal if inhaled, swallowed, or absorbed through the skin. This material causes burns and is extremely destructive to tissue of the mucous membranes and upper respiratory tract, eyes, and skin. Inhalation may be fatal as a result of spasm, inflammation, and edema of the larynx and bronchi, chemical pneumonitis, and pulmonary edema. Symptoms of exposure may include burning sensation, coughing, wheezing, laryngitis, shortness of breath, headache, nausea, and vomiting.

Thallium and Thallium Nitrate: Inhalation of dust produced by thallium and thallium compounds may be irritating to the upper respiratory tract. Repeated or prolonged skin exposure may cause dermatitis. Chronic exposure to the eye may cause conjunctivitis.

Thallium compounds are cumulative poisons. Ingestion of thallium and thallium compounds may cause metallic taste in the mouth, nausea, vomiting, diarrhea, abdominal pain and gastrointestinal hemorrhage which may be delayed for several days followed by circulatory collapse, ptosis, skin eruptions, peripheral neuritis, tremors, and leg pains with weakness. Slowed reflexes, convulsions, disorientation, psychosis, restlessness, slurred speech, liver and kidney damage may also occur. Pulmonary edema, delirium, and coma may precede death in respiratory failure. Rarely, inorganic nitrates may be converted to nitrites by nitrate-reducing bacteria in the digestive tract, resulting in methemoglobinemia. Repeated or prolonged exposure to thallium and thallium compounds may cause hair loss, soreness of the mouth, trembling, severe weight loss, endocrine disorders, and emotional disturbances.

Medical Conditions Generally Aggravated by Exposure: Nitric Acid: eye disorders, skin disorders, respiratory disorders, and allergies

Listed as a Carcinogen/Potential Carcinogen:

	Yes	No
In the National Toxicology Program (NTP) Report on Carcinogens		X
In the International Agency for Research on Cancer (IARC) Monographs		X
By the Occupational Safety and Health Administration (OSHA)		X

EMERGENCY AND FIRST AID PROCEDURES:

Skin Contact: Remove contaminated shoes and clothing. Rinse affected area with large amounts of water followed by washing the area with soap and water. Watch for chemical irritations and treat them accordingly. Obtain medical assistance if necessary.

Eye Contact: Immediately flush eyes, including under the eyelids, with copious amounts of water for at least 15 min. Obtain medical assistance.

Inhalation: If inhaled, move the victim to fresh air. If breathing is difficult, give oxygen; if the victim is not breathing, give artificial respiration. Obtain medical assistance if necessary.

Ingestion: If ingestion occurs, wash out mouth with water. **DO NOT** induce vomiting. Obtain medical assistance immediately.

NOTE (Nitric Acid): Wash affected skin areas with 5 % solution of sodium bicarbonate ($NaHCO_3$). If ingested, the risk versus the benefit of the passage of a naso-gastric tube is debatable. Activated charcoal is of no value. **DO NOT** give the exposed person bicarbonate to neutralize the material.

TARGET ORGAN(S) OF ATTACK: **Nitric Acid:** skin, teeth, eyes, and upper respiratory tract
Thallium and Thallium Nitrate: nervous system

SECTION VII. PRECAUTIONS FOR SAFE HANDLING AND USE

Steps to be Taken in Case Material Is Released or Spilled: Notify safety personnel of spills. Surfaces contaminated with spills should be covered with soda ash or sodium bicarbonate to neutralize the acid. Place the neutralized material into containers suitable for eventual disposal, reclamation, or destruction.

Waste Disposal: Follow all federal, state, and local laws governing disposal.

Handling and Storage: Provide general and local explosion proof ventilation systems to maintain airborne concentrations below the TLV. Provide approved respiratory apparatus for non-routine or emergency use. Use an approved filter and vapor respirator when the vapor or mist concentrations are high. Wear gloves and chemical safety glasses where contact with the liquid or high vapor concentrations may occur. An eye wash station and washing facilities should be readily available near handling and use areas.

> **NOTE:** Contact lenses pose a special problem; soft lenses may absorb irritants and all lenses concentrate them. **DO NOT** wear contact lenses in the laboratory.

Store this material in its original bottle at room temperature. It must be protected from moisture.

SECTION VIII. SOURCE DATA/OTHER COMMENTS

Sources:
MDL Information Systems, Inc., MSDS *Nitric Acid*, 16 September 1999.
MDL Information Systems, Inc., MSDS *Thallium*, 21 March 2000.
MDL Information Systems, Inc., MSDS *Thallium Nitrate*, 2 June 1999.
The Merck Index, 11th Ed., 1989.
The Sigma-Aldrich Library of Chemical Safety Data, Ed. II, 1988.